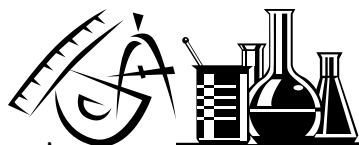


Significant Figures

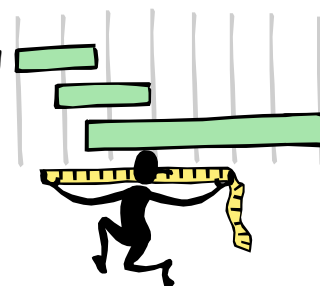
Measuring with Significant Figures

1. Always estimate one place greater than the scale itself reads.
2. That last value is an educated guess.
3. Any values that you can measure are significant.



Identifying Significant Figures

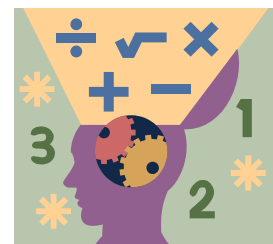
1. All reported non zero digits in a value are significant.
893 cm - 3 sig figs 12.345 in - 5 sig figs
2. All digits are significant if the value ends with a decimal point.
8000. m - 4 sig figs 980. g - 3 sig figs
3. Zeroes to the right of the decimal point and to the right of a significant figure in a value are significant.
5.0 cm - 2 sig figs 0.00341 m - 3 sig figs 500 in - 1 sig fig
4. Zeroes between significant figure in a value are significant.
4002 g - 4 sig figs 3400.0 kg - 5 sig figs



Multiplication/Division with Significant Figures

1. Identify the number of significant figures in each value.
2. Multiply or divide the values
3. The value with the fewest significant figures will determine the number of significant figures in the answer.

42 m (2)	234 m (3)	3780 m (3)
$\times 2 \text{ m (1)}$	$\times 56 \text{ m (2)}$	$\times 45.6 \text{ m (3)}$
84 m ²	13184 m ²	172368 m ²
with sig figs 80 m ² (1)	13000 m ² (2)	172000 m ² (3)



Addition/Subtraction with Significant Figures

1. Line up the decimal point of each value to be added or subtracted.
2. Locate the last significant figure in each value
3. Add or subtract the values
4. The significant figure furthest to the left of the original values will determine the location of the significant figure of the answer.



$$346.1\text{cm} + 45.98\text{cm}=? \quad 65.85\text{cm} + 2100\text{cm}=? \quad 92.6\text{in} - .212\text{in}=?$$

346.1 cm	65.85 cm	92.6 in
$+ 45.98 \text{ cm}$	$+ 2100 \text{ cm}$	$- .212 \text{ in}$
392.08 cm	2165.85 cm	92.388 in
With sig figs 392.1 cm	2200 cm	92.4 in

Significant Figures in Logarithms (pH & pOH)

1. Only those numbers to the right of the decimal place count as significant.
pH=10.26 -2 sig figs and corresponds to a $[H^+] = 5.5 \times 10^{-11}M$
 $[H^+] = 1.9 \times 10^{-9}M$ - 2 sig figs corresponds to a pH=8.72 -2 sig figs